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**Ionic Compound Formula Writing Worksheet**

Write chemical formulas for the compounds in each box. The names are listed by listing the anion before the cation and writing. Example: The first box is the phosphate ion and the second box is the sodium ion, so you should write Na<sub>3</sub>PO<sub>4</sub>.

	tin(II)	tin(IV)	gallium	silver	lead(II)
chloride					
nitrate					
sulfate					
iodide					
acetate					

Write the formulas for the following compounds:

- 1) copper (II) chloride
- 2) lithium nitrate
- 3) vanadium (II) acetate
- 4) manganese (II) nitrate
- 5) barium oxide
- 6) sodium sulfide
- 7) aluminum acetate
- 8) potassium permanganate
- 9) chromium (II) cyanide
- 10) tin (II) iodide
- 11) vanadium (II) fluoride
- 12) ammonium nitrate

**NAMING ACIDS**

Name Key

Name the following acids.

1. HNO<sub>3</sub> nitric acid
2. HCl hydrochloric acid
3. H<sub>2</sub>SO<sub>4</sub> sulfuric acid
4. H<sub>2</sub>SO<sub>3</sub> sulfurous acid
5. HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> acetic acid
6. HBr hydrobromic acid
7. HNO<sub>2</sub> nitrous acid
8. H<sub>3</sub>PO<sub>4</sub> phosphoric acid
9. H<sub>2</sub>S hydrosulfuric acid
10. H<sub>2</sub>CO<sub>3</sub> carbonic acid

Write the formulas of the following acids.

11. sulfuric acid H<sub>2</sub>SO<sub>4</sub>
12. nitric acid HNO<sub>3</sub>
13. hydrochloric acid HCl
14. acetic acid HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> or CH<sub>3</sub>COOH
15. hydrofluoric acid HF
16. phosphorous acid \* H<sub>2</sub>PO<sub>3</sub>
17. carbonic acid H<sub>2</sub>CO<sub>3</sub>
18. nitrous acid HNO<sub>2</sub>
19. phosphoric acid H<sub>3</sub>PO<sub>4</sub>
20. hydrosulfuric acid H<sub>2</sub>S

Name \_\_\_\_\_

**Binary Ionic Compounds - Worksheet**

Which of the following are subscripts? Which are superscripts?

1. O<sub>2</sub> - \_\_\_\_\_
2. S<sub>8</sub> - \_\_\_\_\_
3. F<sup>2-</sup> - \_\_\_\_\_
4. I<sub>3</sub> - \_\_\_\_\_
5. H<sup>3+</sup> - \_\_\_\_\_
6. Si<sub>2</sub> - \_\_\_\_\_
7. Cu<sup>-2</sup> - \_\_\_\_\_
8. Br<sup>-</sup> - \_\_\_\_\_

What would the chemical formula be for the following compounds?

1. Mg<sup>2+</sup> + Br<sup>-</sup> - \_\_\_\_\_
2. Al<sup>3+</sup> + Br<sup>-</sup> - \_\_\_\_\_
3. Na<sup>+</sup> + Cl<sup>-</sup> - \_\_\_\_\_

Name \_\_\_\_\_

**Molecular Polarity and Dipole-Dipole Forces - Worksheet**

In what order would the following elements appear according to dipole-dipole interactions? (REMEMBER: The element that is MORE electronegative goes after the one that is LESS electronegative)

1. H + O \_\_\_\_\_
2. O + C \_\_\_\_\_
3. Br + Mg \_\_\_\_\_

## Binary Ionic Compounds

Name: \_\_\_\_\_ Per: \_\_\_\_\_

**Directions:** First quickly scan the worksheet and circle any metals (as a symbol or as a name) that is a transition metal. Then either give the name or chemical formula as necessary for each problem below.

*Remember that transition metals can have multiple oxidation states, so you are required to indicate the appropriate oxidation state in parenthesis when naming them. You must use the subscript and oxidation state of the anion to help determine the oxidation state of the transition metal.*

- |   |  |
|---|--|
| 1. LiF <u>Lithium Fluoride</u>                                  | 2. lithium chloride <u>LiCl</u>                                |
| 3. Li <sub>2</sub> O <u>Lithium Oxide</u>                       | 4. lithium nitride <u>Li<sub>3</sub>N</u>                      |
| 5. Li <sub>3</sub> P <u>Lithium Phosphide</u>                   | 6. beryllium fluoride <u>BeF<sub>2</sub></u>                   |
| 7. BeO <u>Beryllium Oxide</u>                                   | 8. beryllium sulfide <u>BeS</u>                                |
| 9. BF <sub>3</sub> <u>Boron Fluoride</u>                        | 10. boron chloride <u>BCl<sub>3</sub></u>                      |
| 11. BB <sub>3</sub> <u>Boron Bromide</u>                        | 12. boron oxide <u>B<sub>2</sub>O<sub>3</sub></u>              |
| 13. BN <u>Boron Nitride</u>                                     | 14. sodium fluoride <u>NaF</u>                                 |
| 15. CuF <u>Copper (I) Fluoride</u>                              | 16. copper (II) chloride <u>CuCl<sub>2</sub></u>               |
| 17. Cu <sub>2</sub> O <u>Copper (I) Oxide</u>                   | 18. copper (II) oxide <u>CuO</u>                               |
| 19. Cu <sub>2</sub> N <u>Copper (I) nitride</u>                 | 20. copper (II) nitride <u>Cu<sub>2</sub>O<sub>3</sub></u>     |
| 21. Na <sub>2</sub> O <u>Sodium Oxide</u>                       | 22. sodium nitride <u>Na<sub>3</sub>N</u>                      |
| 23. PbF <sub>2</sub> <u>Lead (II) Fluoride</u>                  | 24. lead (IV) fluoride <u>PbF<sub>4</sub></u>                  |
| 25. PbS <u>Lead (II) Sulfide</u>                                | 26. lead (IV) sulfide <u>PbS<sub>2</sub></u>                   |
| 27. Pb <sub>2</sub> N <sub>4</sub> <u>Lead (IV) Nitride</u>     | 28. lead (IV) oxide <u>PbO<sub>2</sub></u>                     |
| 29. FeF <sub>3</sub> <u>Iron (III) Fluoride</u>                 | 30. iron (II) bromide <u>FeBr<sub>2</sub></u>                  |
| 31. Fe <sub>2</sub> O <sub>3</sub> <u>Iron (III) Oxide</u>      | 32. chromium (VI) phosphide <u>Cr<sub>2</sub>P<sub>4</sub></u> |
| 33. FeP <u>Iron (III) Phosphide</u>                             | 34. aluminum fluoride <u>AlF<sub>3</sub></u>                   |
| 35. AlI <sub>3</sub> <u>Aluminum Iodide</u>                     | 36. aluminum oxide <u>Al<sub>2</sub>O<sub>3</sub></u>          |
| 37. Mn <sub>2</sub> O <sub>7</sub> <u>Manganese (VII) Oxide</u> | 38. manganese (IV) bromide <u>MnBr<sub>4</sub></u>             |
| 39. SnSe <u>Tin (II) Selenide</u>                               | 40. tin (IV) oxide <u>SnO<sub>2</sub></u>                      |

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**Type II Binary Ionic Compounds** contain Transition metals (including the Group III, IV, V, VI metals, except for Al) with non-metal ions.

Show the correct name for the following compounds.

### Give correct formulas for these Type II binary compounds

copper(II) iodide	copper(I) oxide
iron(II) sulfide	manganese(IV) oxide
gold(III) chloride	gold(I) sulfide
silver bromide*	cadmium sulfide*
vanadium(V) nitride	chromium(VI) oxide
lead(IV) sulfide	zinc selenide*
gallium(I) iodide	vanadium(III) chloride
nickel(III) selenide	iron(III) bromide
silver sulfide*	tin(IV) chloride
manganese(II) fluoride	antimony(V) bromide
cobalt(III) oxide	titanium(IV) oxide

### Give correct Type II names for the following compounds

FeCl <sub>3</sub>	Au <sub>2</sub> O
FeCl <sub>2</sub>	AuCl <sub>3</sub>
V <sub>2</sub> O <sub>5</sub>	PbO
MnO <sub>2</sub>	Pb <sub>2</sub> S <sub>3</sub>
NiF <sub>3</sub>	Hg <sub>2</sub> F <sub>2</sub>
CuS	HgI <sub>2</sub>
CuBr	CdSe
Fe <sub>2</sub> O <sub>3</sub>	ZnBr <sub>2</sub>
SnF <sub>4</sub>	ScCl <sub>3</sub>
Ga <sub>2</sub> S	MnCl <sub>2</sub>
AgI	NiCl <sub>2</sub>

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