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Electron Configuration Worksheet

In the space below, write the abbreviated electron configurations of the following elements. Use Figure 11.23 on Pg. 381.

- 1) hydrogen _____
- 2) lithium _____
- 3) sodium _____
- 4) helium _____
- 5) neon _____
- 6) argon _____
- 7) magnesium _____
- 8) carbon _____
- 9) silicon _____
- 10) manganese _____

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2$ _____
- 12) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1 3d^5 4p^5$ _____
- 13) $[Kr] 5s^2 4d^9 5p^1$ _____
- 14) $[Xe] 6s^2 4f^6$ _____
- 15) $[Rn] 7s^1$ _____

Name: _____ Date: _____ Block: _____

Ionic bonding practice

How many valence electrons do each of the following elements have?

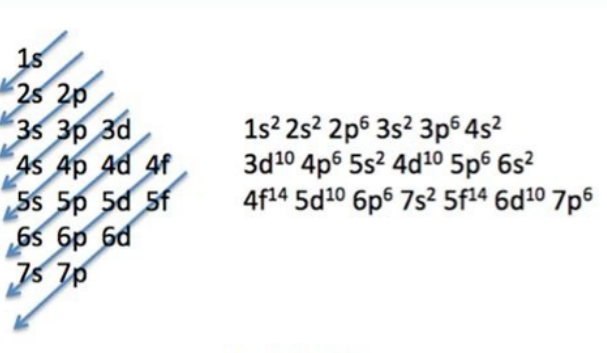
- 1) carbon _____
- 2) selenium _____
- 3) xenon _____
- 4) potassium _____

Which of the following ions are likely to be formed? Circle one:

- | | | | | | |
|-------------|-----|----|---------------|-----|----|
| 5) N^{2+} | yes | no | 8) Al^{2+} | yes | no |
| 6) He^+ | yes | no | 9) P^{-3} | yes | no |
| 7) F^{-1} | yes | no | 10) Mg^{-1} | yes | no |

Write the formula which result from combining the following ions:

- 11) K^+ and Br^- _____
- 12) Li^+ and O^{2-} _____
- 13) Ca^{2+} and Cl^- _____
- 14) Al^{3+} and S^{2-} _____
- 15) Fe^{2+} and S^{2-} _____
- 16) Na^+ and NO_3^- _____
- 17) Na^+ and SO_4^{2-} _____
- 18) NH_4^+ and CO_3^{2-} _____
- 19) Fe^{3+} and CrO_4^{2-} _____
- 20) Al^{3+} and PO_4^{3-} _____



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