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TALLER DE QUIMICA

1) ¿Qué propiedades físicas presentan el agua de acuerdo con su tipo de enlace?

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El agua es uno de los pocos elementos que se presentan en los 3 estados básicos de la materia, es el único que tiene menos densidad en estado sólido que líquido.

2) En un conjunto de moléculas, identifica el tipo de fuerzas intermoleculares que se llevan a cabo en cada caso:

R//

- A. Interacciones dipolo-dipolo
- B. Interacciones dipolo-dipolo
- C. Fuerzas de London
- D. Fuerzas de London
- E. Interacciones dipolo-dipolo

3) Señala cuales de los siguientes compuestos solo tienen fuerza de London: tetra cloruro de carbono (CCl4), penta cloruro de fósforo (PCl5), cloro etano (C2H5Cl), metanol (CH3OH) y cloruro de litio (LiCl)

R//

London: (CCl4), (C2H5Cl), (CH3OH)

4) Describe las fuerzas intermoleculares que se esperan en cada uno de los siguientes compuestos: Oxido sulfuroso (SO2), Dióxido de carbono (CO2) y Metilamina (CH3NH2).

R//

Oxido sulfuroso: dipolo - dipolo cuando una molécula de azufre se une con 2 de oxígeno al tener diferente electronegatividad se presentan fuerzas dipolo.



Enlaces Intermoleculares
 Química Inorgánica

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PRÁCTICA PARA LA CLASE

Conteste las preguntas 1 y 2 de acuerdo a la información de la siguiente tabla:
La tabla presenta la electronegatividad de 4 elementos X, J, Y y L

Elemento	X	J	Y	L
Electronegatividad	4.0	1.5	0.9	1.6

- De acuerdo con la información de la tabla, es válido afirmar que el compuesto con mayor carácter iónico es
a) LX b) JL c) YJ d) YX
- De acuerdo con la información de la tabla, es válido afirmar que el compuesto de mayor carácter covalente es
a) LY b) JL c) YX d) YJ

VALORES DE ELECTRONEGATIVIDAD SEGÚN LA ESCALA DE PAULING DE ALGUNOS ELEMENTOS

H	Li	Be	B	C	N	O	F
0.97	1.0	2.0	2.0	2.5	3.0	3.5	4.0
Na	Mg	Al	Si	P	S	Cl	Br
0.9	1.0	1.5	1.7	2.1	2.5	3.0	2.8
K	Ca	Ga	Ge	As	Se	Br	I
0.8	1.0	1.5	2.0	2.2	2.5	2.8	2.5
Rb	Sr	In	Sn	Sb	Te	I	At
0.8	1.0	1.5	1.7	1.9	2.0	2.2	2.2
Cs	Ba	Tl	Pb	Bi	Po	At	At
0.8	0.9	1.4	1.5	1.7	1.8	1.9	1.9

- Teniendo en cuenta que los valores de la electronegatividad según la escala de Pauling de los elementos siguientes son: H: 2.1 ; O: 3.5 ; Na: 0.9 ; S: 2.5 y Cl: 3.0 ¿Cual de los siguientes enlaces es más polar?
a) H-O b) H-Na c) H-S d) H-Cl
- Los tipos de enlace que pueden darse entre dos átomos pueden ser:
A. Iónico y covalente.
B. Iónico, covalente y metálico.
C. Iónico, covalente, metálico y por fuerzas de Van der Waals.
D. Iónico, covalente, metálico, por fuerzas de Van der Waals y por puente de hidrógeno.
- En las reacciones químicas, las partículas de los átomos que interactúan para producir nuevas sustancias son
A. los electrones que hay en el núcleo.
B. los protones del último nivel de energía.
C. los neutrones de los orbitales enlazados.
D. los electrones de valencia.
- ¿En cual de los compuestos siguientes tiene un enlace fundamentalmente iónico?

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He will recognize that, from the concept of hybridization, you can identify the form of geomism that the spots that are discriminated against if it is a polar mol or not polar, regardless of the type of type of Link that presents and deduces certain fanic properties that show substances. What is the difference between intermolecular and intramolecular forms? In this unit, you have acquired the necessary powers that will allow you to predict the fanic properties present in the substances, depending on the type of link between the other or the compound or the substance. This difference is important, as the terms are similar and can be easily confused. It can be used the mol of water as an example of this type of link. Slideshare uses cookies to improve the functionality and performance of our site to offer relevant advertising. Metary The Methanic Ligament is the result of the forces that occur within the mollets in metallic substances. The force of this relationship will depend on the polarity that the molers are. First, it is worthwhile that intermolecular forms are not more than a quantum way to link. In addition, these compounds are not very solid in water. Since metals are a small number of elf, they end up being great conductors of electricity and tumic. Check our conditions of use and our privacy policy to obtain more information. Therefore, you have reviewed their concepts around the Quantic Link, in addition to understanding that it is based on the foundation that the link is a forgiveness that is maintained, giving, transferring, transferrring, Sharing, sharing, or providing Valáca and Valencia, and in this way, to acquire a similar configuration of a nobleman. At of covalent compounds can be kept together among the same electrostatic interactions of characters that are genus known as "intermolecular forms". This union is triggered by the electrostatic attractive that occurs between the other with different loads. Covalent covalent ligament produces a force that results in the division of pairs of elf in 2 UTS not methalic. Dipolo-Dipolo Interaction in Covalent Jackets There are other Dipol-Dipol. Likewise, you knew how much the classification of the ligament quommic and recognized the existence of the vain between moms and sauce. The springs inserted in the water attract and are also attracted to the surrounding Molens. This is because the mollant, even if it is nonsense, has many elo's trons, which move very approximately. With this, polarized mold can generate the polarization of some molts that are on the side, through an elegric induction. The formation of the phases is due precisely for the forces, when flash or sine occurs, they remain united, due to this phenomenon, the points of ebululicion. This way, these molers end up losing strongly, creating a movie in the superficial layer of the water, as if it were a planned movie. These links are characterized by being stronger, as electronegativity is quite different between the elements. Within each type of van der waals, the intensity increases with the surface of the mol; This is, with size (which in general, although not always, is linked to molecular weight). Now you can see the characteristics of each The types of quamic ligament: Yaian, polar covalent, not polar covalent, coordinated covalent, metallic bridge and hydrogen, supporting you by your representation with the lewis diagram, also The fanic properties presented by the substances derived from the type of link, among them: solubility, fusion point, ebululicion point, elegric conductivity, technique conductivity, hardness, malleability and ductility. In addition, the elegant dipole is not formed. However, when the molers approach, the temporary dipoles are formed by induction. Tweet Share 0 Reddit +1 Pocket LinkedIn 0 Slideshare uses cookies to improve the functionality and performance of our site to offer relevant advertising. If you continue to look for this site, accept the use of cookies. Most covalent compounds have low points of fusion and ebululicion. In water, however, it occurs more intensely. Intermolecular forms are characterized by maintaining a mol © United. DFIE - IPN (2015). In this form of union, the mollets considered to interact polar so that the opposite pills are preserved. This allows a attractive of both to be believed, even if it is one day. In this case, the interaction of hydrogen with oxygen, phlogãnio and nitron. It does not occur. Before deepening the question of the intermolecular forms, it is essential to start the article with the basic concepts to facilitate the understanding of the subject. That is, the forces that unite utomos (athic links) as covalent, larger than the forms that unite moles (intermolecular links). The £ £ £ © equivalent is established between metal and the number of meal through the transfer of elf. The intermolecular van is the union that, as a result of the Electrostatic that is established among the molers, he can keep them together in a crystalline network. This allowed to clarify the main information related to intermolecular forms in order to insert more details about this important question. In this way, the forces eventually balance themselves between being. Intermolecular forms cause different fanic states in quantic compounds. The Hydrode Bridge is not really quantic connections, but a very strong dipole-dipolus interactions. Therefore, there are 3 types: Hydrogen Link - Strong Ligament Intensity; Permanent Dipolo (Dipolo-Dipolo); Intermediate Intensity; INDUCED DIPOLO (London Force); BILL LINK INTENSITY; Hydrogen League Hydro-gereques can be observed in mollets whose hydrogen is attached to electronegative elements with low, such as phlogã, nitregãnio and oxygen. In the Dipolo-Dipolo, the asymetrically distributed ENTRICAL ELABONS, so that the most electronegative elements end up attracting elors. Both dipoles attract. Hydrogen ligales in the water generate high superficial tension. In this case, the trons are Smister distributed. Therefore, they are related to the quantic links, which are the purpose of making molers join or separate. The NaCl, Cabr2 and K2SO4 salts are an lanic vain, their mansions are discret in a state, their vain is strong and almost all tall points. It is possible to see that throughout the Lyh, the phenomenon described above can be observed. Anything that does not fit this situation is considered an intramolecular force. The intermolecular forms observed in the induced dipole are 10 times more than the permanent dipole connections. Classification is important for Intermolecular can be classified according to its intensity. These forces are classified into two basic types: Hydge Bridge Links and Van Der Waals forms. This Mollant will polarize for a few moments. This superficial tension can explain certain feats. The existence of the mollant is due to these forms; For example, it is necessary 920 kJ of energy to decompose a water vapor mole in hydrãgen (h) and oxygen (o), which indicates the existing intermolecular forction (quarmic connection). But only 40.7 kJ are necessary to convert a water mole into no vapor of water vapor at 100 ° C (in this case, these intermolecular forms are reflected). Although there are different types of intermolecular forms, as shown below, all of them are much less forms than an ibanico or covalent ligation, f hydrãgen link: Link that is established between a time of a mol Å © Cup, which is being linked to F, or ã "ã (the 3 electronegative volumes) has a large positive charge density, and one input of F, or another of another mol that will have a great density Negative Cargo: Van der Waals: Induced (Apolar-Apolar) Cloud (Apolar) and form a instant dipole, which induces another dipole in a nearby mole. Intermolecular Force or Van der Waals [Image of Prar. The induced dipole or the forms of London the induced dipole occurs when all the polar or nonpolar slaughter, which are attracted to -gravitational. Check our privacy and our use conditions for more information. Some properties such as the point of fusion and ebululicion of covalent compounds can be explained based on intermolecular forms that maintain united molecase, these forces are days; Therefore, in general, it is necessary to provide less energy for substances to change the status of aggregation. The hydrity bridge is presented in the compounds whose covalent connection is contained in hydro -dipole. Thus, it may happen that at one point a mol © Crows have more castles on one side. For example, insects can walk on the water and this is because the drops of water are a spray form. To find out how these attractions are generated, click on each image: among the utoms, the substances, there are forces that are attracted, they are called Intermolecular, in some days and in other intermediate or strong. On the other hand, the springs located in the surface of the lichia are attracted to the jumps located below or next to them. Permanent dipole (dipole) the permanent dipole (dipole-dipole) is the intermediate intermediate forms observed between the gluers of the polar compounds. While intermolecular forms occur between the springs, the intramolecular is observed within these same mollets. Cloud Electronics Distortion between two different electronegativity [image of your own elaboration] to understand the quantum vain that you must trust in Varia Stratão © didicos, as the knowledge of the elements in the elements dangerous table in which the trons of elors are found that they are not in their last, They are Valencia's ELABRES and the links of ligament that, with their sounds, as well as through the Lewis diagram, it is possible to understand the way the elements are linked to constitute the mollets, thus forming different quantic links (lanic, covalent, metallic bridge and hydrity) and, so much hybridize, they are different forms in the space; Therefore, they have polar and not polar moldings, which results in different properties, for example: ctric, rumical conductivity, solubility, fuse point and ebululicion point, which is perfectly related to the link and ã link Hybridization that provides the form of the mol © in the space. Intramolecular forms can be classified as: ionic the ligament £ Force.

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